Quiz 4 Key / Fall 2017

**1.**

5d

4f

6s

5p

4d

5s

4p

3d

4s

3p

3s

2p

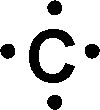
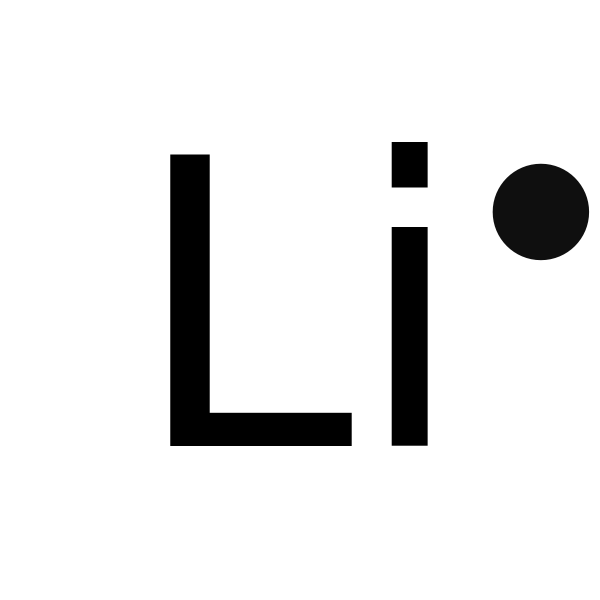
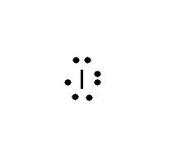
2s

1s

**2.** 1s22s22p63s23p64s23d104p65s1

**3.** As

**4.** Ionization energy (IE) of the elements within a period generally increases from left to right on the periodic table. This is due to valence shell stability. IE generally decrease from top to bottom on the periodic table. This is due to electron shielding.

**5.** 

**6. False**

**7. False**